Name:		Roll No.		Subject	Chemistry
Test Type	Only Exercise MCQs	Class	11th	Date	
Chapter	01	Unit	01	Time	

Q. No. 1: Multiple Choice Questions (MCQs)

I. Which scientist first observed the periodicity in the elements?

a) J. Newlands b) L. Meyer c) J. W. Dobereiner d) D. I. Mendeleev

II. Identify the element which has three electron shells, belongs to the "s" block, and has two electrons in its outermost shell.

a) Calcium b) Sodium c) Magnesium d) Potassium

III. Which one of the following statements is correct about metallic character?

a) It decreases from top to bottom in a group. b) It increases from top to bottom in a group.

c) It remains constant from left to right in a period. d) It increases from left to right in a period.

IV. Which property increases as you go down a group in the periodic table?

a) Atomic radius b) Electron affinity c) Electronegativity d) Ionization energy

V. Which set of conditions results in higher ionization energy?

a) Smaller atom and greater nuclear charge b) Smaller atom and smaller nuclear charge

c) Larger atom and greater nuclear charge d) Larger atom and smaller nuclear charge

VI. Which of the following atoms shows more than one (variable) oxidation state?

a) Sodium b) Magnesium c) Aluminum d) Phosphorus

VII. Which is the correct general trend for the variation of electron affinity in a group?

a) It becomes less negative from top to bottom. b) It becomes more negative from top to bottom.

c) It remains the same. d) It has no definite trend and changes irregularly.

VIII. What is the oxidation state of sulfur in the sulfate ion (SO_4^{2-})?

a) +4 b) +2 c) +6 d) 0

IX. What is the correct trend for the variation of electronegativity across a period in the periodic table?

a) It decreases from left to right across a period. b) It increases from left to right across a period.

c) It remains constant. d) It has no definite trend.

X. The atomic radius generally _____ across a period in the periodic table.

a) Increases b) Decreases c) Remains constant d) First increases, then decreases

XI. Which one of the following elements has the highest ionization energy?

a) Sodium (Na) b) Magnesium (Mg) c) Aluminum (Al) d) Argon (Ar)